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Equilibrium Constant (p.429 in the
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Section 8.5: Calculations Involving Basic Solutions

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= 12.00 mmol - 5.40 mmol Unreacted ethanoic acid = 6.60 mmol Since 18.00 mL of NaOH(aq) was added to 20.00 mL of ethanoic acid solution, the total volume is now 38.00 mL. Solution: The concentration of unreacted ethanoic acid, [HC₂H₃O₂(aq)], can be determined: $\frac{23.2 \text{ HC}_2\text{H}_3\text{O}_2(\text{aq})}{23.2 \text{ HC}_2\text{H}_3\text{O}_2(\text{aq}) + 23.2 \text{ [HC}_2\text{H}_3\text{O}_2(\text{aq})] (6.60 \text{ mmol})} = 38.00 \text{ mL}$

Section 8.7: Acid-Base Titration Tutorial 1 Practice, page 547

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